

**CHEMISTRY TEST PAPER WITH ANSWER & SOLUTIONS**  
**FINAL NEET(UG)-2023 (EXAMINATION)**  
**Held On Sunday 7th MAY, 2023**  
**CODE - F6**

**Chemistry : Section-A (Q. No. 051 to 085)**

**51.** Given below are two statements : one is labelled as **Assertion A** and the other is labelled as **Reason R** :

**Assertion A** : Metallic sodium dissolves in liquid ammonia giving a deep blue solution, which is paramagnetic.

**Reason R** : The deep blue solution is due to the formation of amide.

In the light of the above statements, choose the **correct** answer from the options given below :

- (1) Both **A** and **R** are true but **R** is **NOT** the correct explanation of **A**.
- (2) **A** is true but **R** is false
- (3) **A** is false but **R** is true
- (4) Both **A** and **R** are true and **R** is the correct explanation of **A**.

**Ans. (2)**

**Sol.** Assertion is correct because all Alkali metals gives deep blue solution by giving electrons.

**Reason** : is incorrect because deep blue solution appears due to the presence of ammoniated electron or solvated electrons.

**52.** The conductivity of centimolar solution of KCl at 25°C is  $0.0210 \text{ ohm}^{-1} \text{ cm}^{-1}$  and the resistance of the cell containing the solution at 25°C is 60 ohm. The value of cell constant is -

- (1)  $3.28 \text{ cm}^{-1}$
- (2)  $1.26 \text{ cm}^{-1}$
- (3)  $3.34 \text{ cm}^{-1}$
- (4)  $1.34 \text{ cm}^{-1}$

**Ans. (2)**

**Sol.** Centimolar solution =  $\frac{1}{100} \text{ M} = 0.01 \text{ M}$

Conductivity ( $k$ ) =  $0.0210 \text{ ohm}^{-1} \text{ cm}^{-1}$

Resistance ( $R$ ) = 60 ohm

$$k = \frac{1}{R} \left( \frac{\ell}{A} \right)$$

$$\Rightarrow 0.0210 = \frac{1}{60} \left( \frac{\ell}{A} \right) \Rightarrow \frac{\ell}{A} = 1.26 \text{ cm}^{-1}$$

**53.** For a certain reaction, the rate =  $k [A]^2 [B]$ , when the initial concentration of A is tripled keeping concentration of B constant, the initial rate would

- (1) increase by a factor of six
- (2) increase by a factor of nine
- (3) increase by a factor of three
- (4) decrease by a factor of nine

**Ans. (2)**

**Sol.** Rate =  $k [A]^2 [B]$

If [A] is tripled and [B] is kept constant.

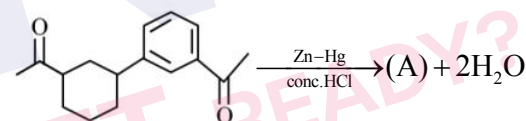
$$r^1 = k [3A]^2 [B]$$

$$r^1 = 9k [A]^2 [B]$$

$$r^1 = 9r$$

Increased by a factor of nine

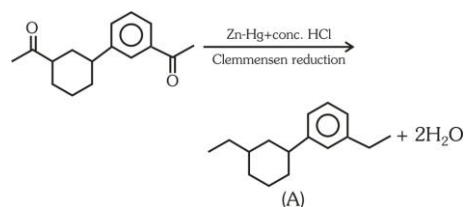
**54.** Identify product (A) is the following reaction :



- (1)
- (2)
- (3)
- (4)

**Ans. (4)**

**Sol.**



55. Which one is an example of heterogenous catalysis ?
- (1) Hydrolysis of sugar catalysed by  $H^+$  ions.
  - (2) Decomposition of ozone in presence of nitrogen monoxide.
  - (3) Combination between dinitrogen and dihydrogen to form ammonia in the presence of finely divided iron.
  - (4) Oxidation of sulphur dioxide into sulphur trioxide in the presence of oxides of nitrogen.

**Ans. (3)**

**Sol.**

- (1)  $C_{12}H_{22}O_{11(aq)} + H_2O \xrightarrow{H^+} C_6H_{12}O_{6(aq)} + C_6H_{12}O_{6(aq)}$   
(Homogeneous reaction)
- (2)  $2O_{3(g)} \xrightarrow{NO(g)} 3O_{2(g)}$  (Homogeneous reaction)
- (3)  $N_{2(g)} + 3H_{2(g)} \xrightarrow{Fe(s)} 2NH_{3(g)}$   
(Reactants and catalyst are in different phase)  
It is heterogeneous reaction
- (4)  $SO_{2(g)} + \frac{1}{2}O_{2(g)} \xrightarrow{NO(g)} SO_{3(g)}$

56. Given below are two statements : one is labelled as **Assertion A** and the other is labelled as **Reason R**.  
**Assertion A** : Helium is used to dilute oxygen in diving apparatus.

**Reasons R** : Helium has high solubility in  $O_2$ .

In the light of the above statements, choose the **correct** answer from the options given below :

- (1) Both **A** and **R** are true but **R** is **NOT** the correct explanation of **A**.
- (2) **A** is true but **R** is false
- (3) **A** is false but **R** is true
- (4) Both **A** and **R** are true and **R** is the correct explanation of **A**.

**Ans. (2)**

**Sol.** Assertion is true because He has low solubility in blood. (NCERT)

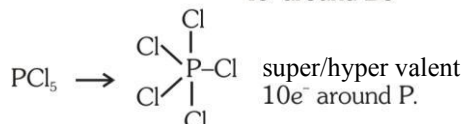
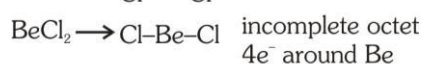
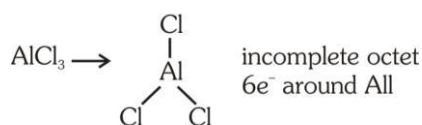
57. Amongst the following, the total number of species NOT having eight electrons around central atom in its outer most shell, is

$NH_3$ ,  $AlCl_3$ ,  $BeCl_2$ ,  $CCl_4$ ,  $PCl_5$  :

- (1) 2
- (2) 4
- (3) 1
- (4) 3

**Ans. (4)**

**Sol.** Total number of species = 3

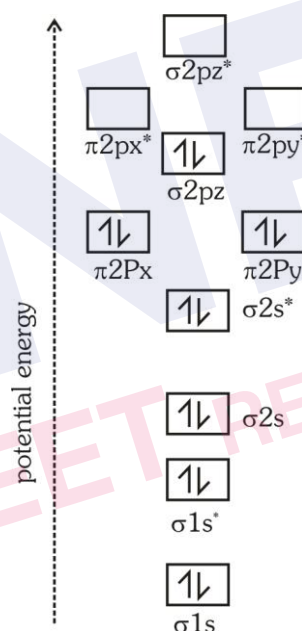


58. The **correct** order of energies of molecular orbitals of  $N_2$  molecule, is

- (1)  $\sigma 1s < \sigma^* 1s < \sigma 2s < \sigma^* 2s < \sigma 2p_z < (\pi 2p_x = \pi 2p_y) < (\pi^* 2p_x = \pi^* 2p_y) < \sigma^* 2p_z$
- (2)  $\sigma 1s < \sigma^* 1s < \sigma 2s < \sigma^* 2s < \sigma 2p_z < \sigma^* 2p_z < (\pi 2p_x = \pi 2p_y) < (\pi^* 2p_x = \pi^* 2p_y)$
- (3)  $\sigma 1s < \sigma^* 1s < \sigma 2s < \sigma^* 2s < (\pi 2p_x = \pi 2p_y) < (\pi^* 2p_x = \pi^* 2p_y) < \sigma 2p_z < \sigma^* 2p_z$
- (4)  $\sigma 1s < \sigma^* 1s < \sigma 2s < \sigma^* 2s < (\pi 2p_x = \pi 2p_y) < \sigma 2p_z < (\pi^* 2p_x = \pi^* 2p_y) < \sigma^* 2p_z$

**Ans. (4)**

**Sol.** Molecular orbital (energy) diagram / sequence of  $N_2$



59. Match **List-I** with **List-II**.

**List-I**

- A. Coke
- B. Diamond
- C. Fullerene
- D. Graphite

**List-II**

- I. Carbon atoms are  $sp^3$  hybridised
- II. Used as a dry lubricant
- III. Used as a reducing agent
- IV. Cage like molecules

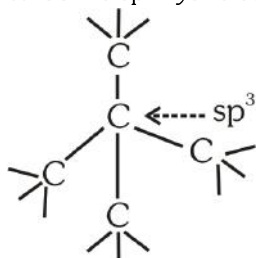
Choose the **correct** answer from the options given below:

- (1) A-IV, B-I, C-II, D-III
- (2) A-III, B-I, C-IV, D-II
- (3) A-III, B-IV, C-I, D-II
- (4) A-II, B-IV, C-I, D-III

**Ans. (2)**

**Sol. Coke** : It is used as reducing agent in carbon reduction methods. (in metallurgical process)

**Diamond** : It is a allotrope of carbon in which each carbon is  $sp^3$  hybridised.



**Fullerene** : It contains pentagonal & hexagonal rings (cage like structure)

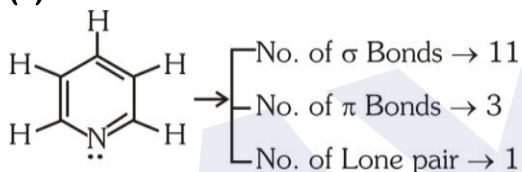
**Graphite** : It is soft solid because graphite layers are bonded with weak Vander Wall attractions.

**60.** The number of  $\sigma$  bonds,  $\pi$  bonds and lone pair of electrons in pyridine, respectively are :

- (1) 12, 3, 0                      (2) 11, 3, 1  
(3) 12, 2, 1                      (4) 11, 2, 0

**Ans. (2)**

**Sol.**



**61.** The element expected to form largest ion to achieve the nearest noble gas configuration is

- (1) F                                  (2) N  
(3) Na                                (4) O

**Ans. (2)**

**Sol.**  $F^{-1}$ ,  $N^{-3}$ ,  $Na^{+}$  &  $O^{-2}$

all ions are isoelectronic containing  $10 e^{-}$

$Z_{eff} \rightarrow Na^{+} > F^{-} > O^{-2} > N^{-3}$

order of radius  $\rightarrow N^{-3} > O^{-2} > F^{-} > Na^{+}$

$\rightarrow$  Nitrogen to achieve Noble gas configuration it gain  $3 e^{-}$ , & form  $N^{-3}$

**62.** Given below are two statements : one is labelled as **Assertion A** and the other is labelled as **Reason R**.

**Assertion A** : A reaction can have zero activation energy.

**Reasons R** : The minimum extra amount of energy absorbed by reactant molecules so that their energy becomes equal to threshold value, is called activation energy.

In the light of the above statements, choose the **correct** answer from the options given below :

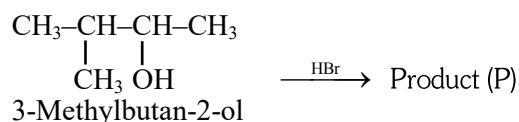
- (1) Both **A** and **R** are true but **R** is **NOT** the correct explanation of **A**.  
(2) **A** is true but **R** is false  
(3) **A** is false but **R** is true  
(4) Both **A** and **R** are true and **R** is the correct explanation of **A**.

**Ans. (3)**

**Sol.** A reaction cannot have zero activation energy.

$E_a$  is minimum extra amount of energy absorbed by reactant molecules so that their energy becomes equal to threshold value.

**63.** Consider the following reaction and identify the product (P).



(1)  $\text{CH}_3\text{CH}=\text{CH}-\text{CH}_3$

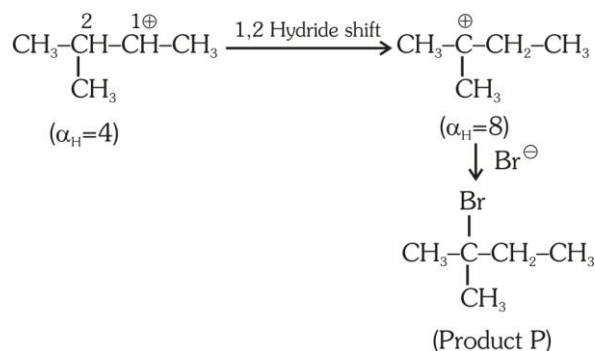
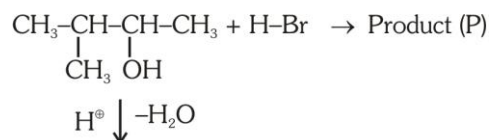
(2)  $\begin{array}{c} \text{CH}_3-\text{CH}-\text{CH}-\text{CH}_3 \\ | \quad | \\ \text{CH}_3 \quad \text{Br} \end{array}$

(3)  $\begin{array}{c} \text{CH}_3 \\ | \\ \text{CH}_3-\text{C}-\text{CH}_2\text{Br} \\ | \\ \text{CH}_3 \end{array}$

(4)  $\begin{array}{c} \text{Br} \\ | \\ \text{CH}_3-\text{C}-\text{CH}_2-\text{CH}_3 \\ | \\ \text{CH}_3 \end{array}$

**Ans. (4)**

**Sol.**



**64.** Given below are two statements : one is labelled as **Assertion A** and the other is labelled as **Reason R** :  
**Assertion A** : In equation  $\Delta_r G = -nFE_{\text{cell}}$ , value of  $\Delta_r G$  depends on  $n$ .

**Reasons R** :  $E_{\text{cell}}$  is an intensive property and  $\Delta_r G$  is an extensive property.

In the light of the above statements, choose the **correct** answer from the options given below :

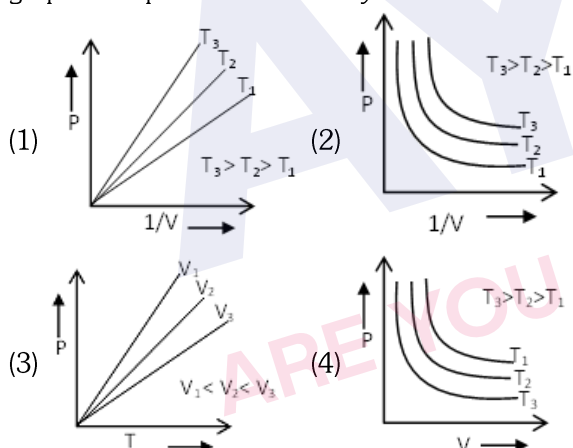
- (1) Both **A** and **R** are true and **R** is **NOT** the correct explanation of **A**.
- (2) **A** is true but **R** is false
- (3) **A** is false but **R** is true
- (4) Both **A** and **R** are true and **R** is the correct explanation of **A**.

**Ans. (4)**

**Sol.**  $\Delta_r G = -nFE_{\text{cell}}$

$E_{\text{cell}}$  is an intensive property and  $\Delta_r G$  is an extensive property as it depends on number of  $e^-$  transferred in cell reaction

**65.** Which amongst the following options is correct graphical representation of Boyle's Law?



**Ans. (1)**

**Sol.** Boyle's law is defined at constant temperature for an ideal gas.

$$P \propto \frac{1}{V}$$

$$P = (nRT) \left( \frac{1}{V} \right) \text{ [straight line equation]}$$

slope of  $P$  versus  $\frac{1}{V}$  curve is  $nRT$

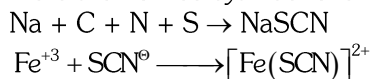
$$\Rightarrow \text{Slope} \uparrow \Rightarrow T \uparrow \therefore T_3 > T_2 > T_1$$

**66.** In Lassaigne's extract of an organic compound, both nitrogen and sulphur are present, which gives blood red colour with  $\text{Fe}^{3+}$  due to the formation of-

- (1)  $\text{NaSCN}$
- (2)  $[\text{Fe}(\text{CN})_5 \text{NOS}]^{4-}$
- (3)  $[\text{Fe}(\text{SCN})]^{2+}$
- (4)  $\text{Fe}_4[\text{Fe}(\text{CN})_6]_3 \cdot x\text{H}_2\text{O}$

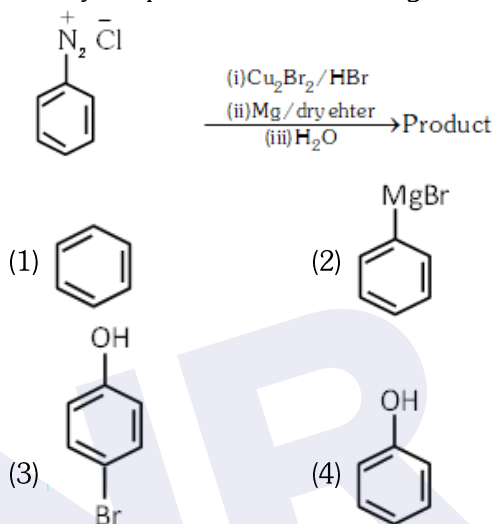
**Ans. (3)**

**Sol.** In case nitrogen and sulphur both are present in an organic compound, sodium thiocyanate is formed, it gives blood red colour and no prussian blue since there are no free cyanide ions



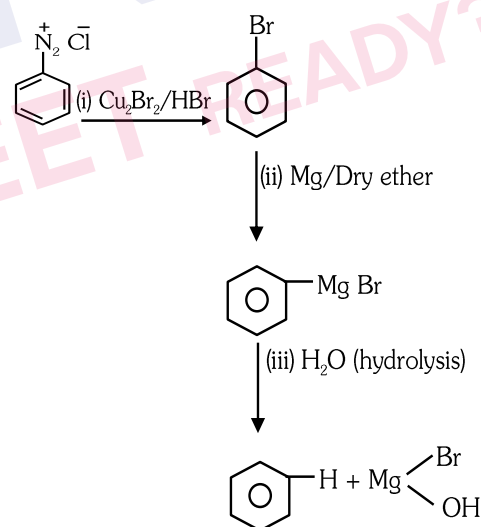
Blood red

**67.** Identify the product in the following reaction :



**Ans. (1)**

**Sol.**



**68.** Select the **correct** Statements from the following :

- A. Atoms of all elements are composed of two fundamental particles.
- B. The mass of the electron is  $9.10939 \times 10^{-31}$  kg.
- C. All the isotopes of a given element show same chemical properties.
- D. Protons and electrons are collectively known as nucleons.
- E. Dalton's atomic theory, regarded the atom as an ultimate particle of matter.

Choose the **correct** answer from the options given below.

- (1) C,D and E only
- (2) A and E only
- (3) B,C and E only
- (4) A,B and C only

**Ans. (3)**

**Sol.** It is statement based question.

Statements B, C & E are correct.

(B) Mass of the electron is  $9.10939 \times 10^{-31}$  kg

(C) All the isotopes of given elements show same chemical properties.

(E) Dalton's atomic theory, regarded the atom as an ultimate particle of matter.

**69.** A compound is formed by two elements A and B. The elements B forms cubic close packed structure and atoms of A occupy  $1/3$  of tetrahedral voids. If the formula of the compound is  $A_xB_y$ , then the value of  $x + y$  is in option

- (1) 4 (2) 3  
(3) 2 (4) 5

**Ans. (4)**

**Sol.**

<b>A</b>	<b>B</b>
$\frac{1}{3} \text{THV}$	CCP
$\Rightarrow Z_A = \frac{1}{3} \times 8 = \frac{8}{3}$	$Z_B = 4$
$\Rightarrow = \frac{8}{3} : 4$	
$\Rightarrow \frac{2}{3} : 1$	
<b>2 : 3</b>	

simplest formula  $\begin{matrix} A_2 & B_3 \\ \downarrow & \downarrow \\ x & y \end{matrix}$

$$x + y = 5$$

**70.** Given below are two statements:

**Statement I :** A unit formed by the attachment of a base to 1' position of sugar is known as nucleoside

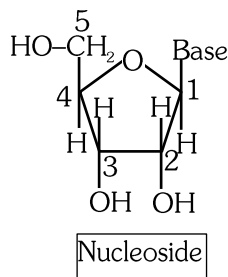
**Statement II :** When nucleoside is linked to phosphorous acid at 5'-position of sugar moiety, we get nucleotide.

In the light of the above statements, choose the **correct** answer from the options given below:

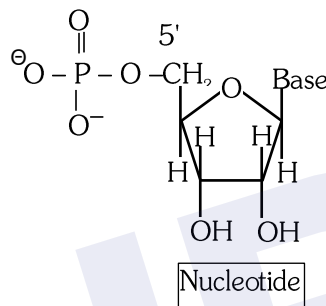
- (1) Both Statement **I** and Statement **II** are false  
(2) Statement **I** is true but Statement **II** is false  
(3) Statement **I** is false but Statement **II** is true  
(4) Both Statement **I** and Statement **II** are true

**Ans. (2)**

**Sol.**



Base link with 1' position of sugar in nucleoside so statement I is correct



→ When nucleoside is linked to **phosphoric acid** at 5' position of sugar moiety we get a Nucleotide

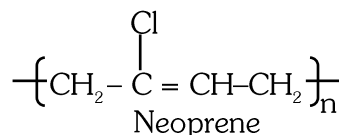
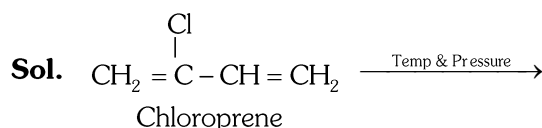


Statement II is Incorrect because not link with **phosphorous acid**.

**71.** Which amongst the following molecules on polymerization produces neoprene?

- (1)  $\begin{matrix} \text{Cl} \\ | \\ \text{H}_2\text{C} = \text{C} - \text{CH} = \text{CH}_2 \end{matrix}$   
(2)  $\text{H}_2\text{C} = \text{CH} - \text{C} \equiv \text{CH}$   
(3)  $\begin{matrix} \text{CH}_3 \\ | \\ \text{H}_2\text{C} = \text{C} - \text{CH} = \text{CH}_2 \end{matrix}$   
(4)  $\text{H}_2\text{C} = \text{CH} - \text{CH} = \text{CH}_2$

**Ans. (1)**



**72.** Taking stability as the factor, which one of the following represents **correct** relationship?

- (1)  $\text{InI}_3 > \text{InI}$  (2)  $\text{AlCl} > \text{AlCl}_3$   
(3)  $\text{TlI} > \text{TlI}_3$  (4)  $\text{TlCl}_3 > \text{TlCl}$

**Ans. (3)**



**Sol.**  $Tl^+$  &  $I^- > Tl^{+3}$  &  $3I^-$

due to inert pair effect  $Tl^+$  is more stable than  $Tl^{+3}$ .

**73.** Some tranquilizers are listed below. Which one from the following belongs to barbiturates?

- (1) Meprobamate (2) Valium  
(3) Veronal (4) Chlordiazepoxide

**Ans. (3)**

**Sol.** Veronal is an example of barbiturates.

**74.** Which of the following statements are **NOT** correct?

- A. Hydrogen is used to reduce heavy metal oxides to metals.  
B. Heavy water is used to study reaction mechanism.  
C. Hydrogen is used to make saturated fats from oils  
D. The H-H bond dissociation enthalpy is lowest as compared to a single bond between two atoms of any element.  
E. Hydrogen reduces oxides of metals that are more active than iron.

Choose the **most appropriate** answer from the options given below:

- (1) B,D only (2) D,E only  
(3) A,B,C only (4) B,C,D,E only

**Ans. (2)**

**Sol.** (D, E) explanation

(D) H-H bond strength/ bond dissociation energy/bond energy of  $H_2$  can not be lowest because bond formed between hydrogen atoms is due to overlapping of  $1s-1s$ .

(E) Hydrogen can not reduces oxides of highly reactive metal.

**75.** Intermolecular forces are forces of attraction and repulsion between interacting particles that will include:

- A. dipole - dipole forces.  
B. dipole - induced dipole forces  
C. hydrogen bonding  
D. covalent bonding  
E. dispersion forces

Choose the **most appropriate** answer from the options given below :

- (1) A,B,C,D are correct (2) A,B,C,E are correct  
(3) A,C,D,E are correct (4) B,C,D,E are correct

**Ans. (2)**

**Sol.** Intermolecular forces means force of attraction between two or more molecules  
dipole-dipole (attraction between two or more polar molecules).

Dipole induced dipole (attraction between polar and non polar molecules)

Hydrogen bonding (it is a special type of dipole-dipole and ion-dipole attraction)

Dispersion forces (mainly acts between non polar molecules).

Covalent bonding (acts between atom not between molecules)

**76.** Amongst the given options which of the following molecules/ion acts as a Lewis acid?

- (1)  $H_2O$  (2)  $BF_3$   
(3)  $OH^-$  (4)  $NH_3$

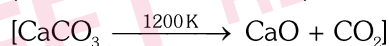
**Ans. (2)**

**Sol.**  $\left. \begin{matrix} H_2O \\ OH^- \\ NH_3 \end{matrix} \right\}$  can not act as lewis acid because they

does not contain vacant orbital

$BF_3 \rightarrow$  Contains vacant orbital on central atom (Boron).

**77.** The **right** option for the mass of  $CO_2$  produced by heating 20 g of 20% pure limestone is (Atomic mass of Ca = 40)



- (1) 1.76 g (2) 2.64 g (3) 1.32 g (4) 1.12 g

**Ans. (1)**

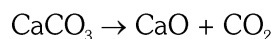
**Sol.** Weight of impure limestone = 20 g

Weight of pure limestone ( $CaCO_3$ ) = 20% of 20 g

$$= \frac{20}{100} \times 20$$

$$= 4g$$

$$n_{CaCO_3} = \frac{4}{100} = 0.04$$



$$n=0.04 \quad n=0.04$$

$$n_{CO_2} = 0.04$$

$$W_{CO_2} = 0.04 \times 44$$

$$= 1.76 g$$

**78.** The relation between  $n_m$ , ( $n_m$  = the number of permissible values of magnetic quantum number (m)) for a given value of azimuthal quantum number (l), is

- (1)  $l = 2n_m + 1$  (2)  $n_m = 2l^2 + 1$   
(3)  $n_m = l + 2$  (4)  $l = \frac{n_m - 1}{2}$

**Ans. (4)**

7

**84.** Homoleptic complex from the following complexes is :

- (1) Diamminechloridonitrito-N-platinum (II)
- (2) Pentaamminecarbonatocobalt (III) chloride
- (3) Triamminetriaquachromium (III) chloride
- (4) Potassium trioxalatoaluminate (III)

**Ans. (4)**

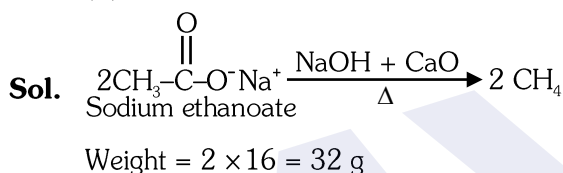
- Sol.** (1)  $[\text{Pt}(\text{NH}_3)_2\text{Cl}(\text{NO}_2)]$   
 (2)  $[\text{Co}(\text{NH}_3)_5(\text{CO}_3)]\text{Cl}$   
 (3)  $[\text{Cr}(\text{NH}_3)_3(\text{H}_2\text{O})_3]\text{Cl}_3$   
 (4)  $\text{K}_3[\text{Al}(\text{C}_2\text{O}_4)_3]$

Option 4 contain all ligands are of same type i.e. why complex will be homoleptic.

**85.** Weight (g) of two moles of the organic compound, which is obtained by heating sodium ethanoate with sodium hydroxide in presence of calcium oxide is :

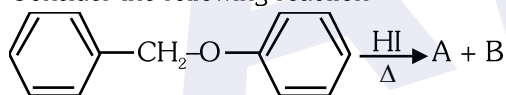
- (1) 32
- (2) 30
- (3) 18
- (4) 16

**Ans. (1)**



**Chemistry : Section-B (Q. No. 086 to 100)**

**86.** Consider the following reaction

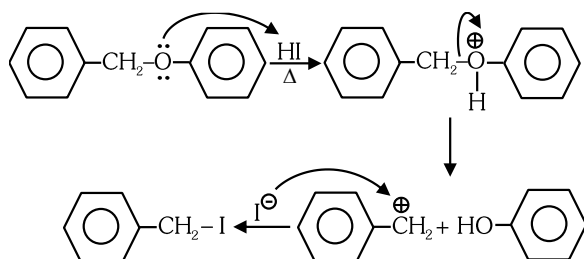


Identify products A and B :-

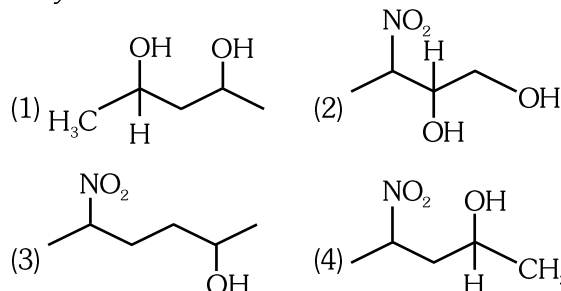
- (1) A =  $\text{C}_6\text{H}_5\text{CH}_2\text{OH}$  and B =  $\text{C}_6\text{H}_5\text{I}$
- (2) A =  $\text{C}_6\text{H}_5\text{CH}_2\text{I}$  and B =  $\text{C}_6\text{H}_5\text{OH}$
- (3) A =  $\text{C}_6\text{H}_5\text{CH}_3$  and B =  $\text{C}_6\text{H}_5\text{I}$
- (4) A =  $\text{C}_6\text{H}_5\text{CH}_3$  and B =  $\text{C}_6\text{H}_5\text{OH}$

**Ans. (2)**

**Sol.**

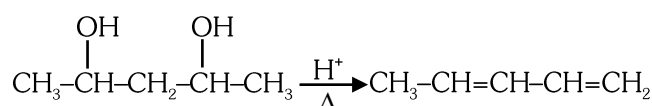


**87.** Which amongst the following will be most readily dehydrated under acidic conditions ?



**Ans. (1)**

**Sol.** Due to presence of conjugation in product.



**88.** The equilibrium concentrations of the species in the reaction  $\text{A} + \text{B} \rightleftharpoons \text{C} + \text{D}$  are 2, 3, 10 and 6 mol  $\text{L}^{-1}$ , respectively at 300 K.  $\Delta G^\circ$  for the reaction is ( $R = 2$  cal/mol K)

- (1) -137.26 cal
- (2) -1381.80 cal
- (3) -13.73 cal
- (4) 1372.60 cal

**Ans. (2)**

**Sol.**  $\text{A} + \text{B} \rightleftharpoons \text{C} + \text{D}$

$$[\text{A}] = 2 \text{ mol L}^{-1}$$

$$[\text{B}] = 3 \text{ mol L}^{-1}$$

$$[\text{C}] = 10 \text{ mol L}^{-1}$$

$$[\text{D}] = 6 \text{ mol L}^{-1}$$

$$\Delta G^\circ = -2.303 RT \log K_{eq}$$

$$= -2.303 RT \log \frac{[\text{C}][\text{D}]}{[\text{A}][\text{B}]}$$

$$= -2.303 \times 2 \times 300 \times \log \frac{10 \times 6}{2 \times 3}$$

$$= -2.303 \times 2 \times 300 \times \log 10$$

$$= -1381.8 \text{ cal}$$

**89.** Given below are two statements :

**Statement I :** The nutrient deficient water bodies lead to eutrophication.

**Statement II :** Eutrophication leads to decrease in the level of oxygen in the water bodies.

In the light of the above statements, choose the **correct** answer from the options given below :

- (1) Both **Statement I** and **Statement II** are false
- (2) **Statement I** is correct but **Statement II** is false.
- (3) **Statement I** is incorrect but **Statement II** is true.
- (4) Both **Statement I** and **Statement II** are true.

**Ans. (3)**

**Sol.** Nutrient enriched water bodies lead to eutrophication.



**90.** Which amongst the following options is the **correct** relation between change in enthalpy and change in internal energy?

- (1)  $\Delta H = \Delta U + \Delta n_g RT$  (2)  $\Delta H - \Delta U = - \Delta n_g RT$   
(3)  $\Delta H + \Delta U = \Delta n_g RT$  (4)  $\Delta H = \Delta U - \Delta n_g RT$

**Ans. (1)**

**Sol.**  $\Delta H = \Delta U + \Delta n_g RT$

**91.** Match **List-I** with **List-II** :

**List-I**

(Oxoacids of

A. Peroxodisulphuric acid

B. Sulphuric acid

C. Pyrosulphuric acid

D. Sulphurous acid

**List-II**

(Sulphur) (Bonds)

I. Two S-OH, Four S=O, One S-O-S

II. Two S-OH, One S=O

III. Two S-OH, Four S=O, One S-O-O-S

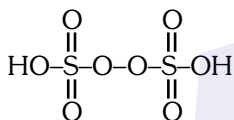
IV. Two S-OH, Two S=O

Choose the **correct** answer from the options given below:

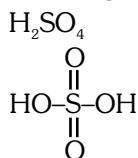
- (1) A-III, B-IV, C-I, D-II (2) A-I, B-III, C-IV, D-II  
(3) A-III, B-IV, C-II, D-I (4) A-I, B-III, C-II, D-IV

**Ans. (1)**

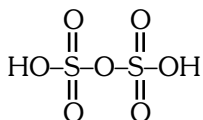
**Sol.** A → Peroxodisulphuric acid



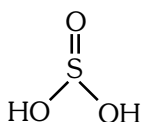
B → Sulphuric acid



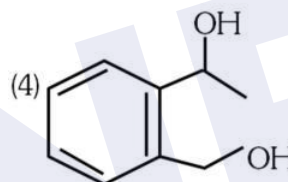
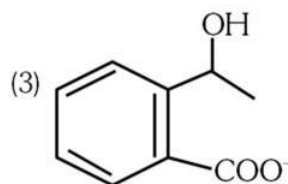
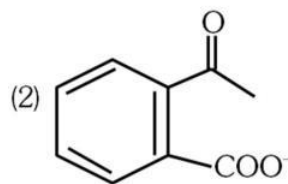
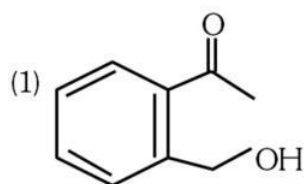
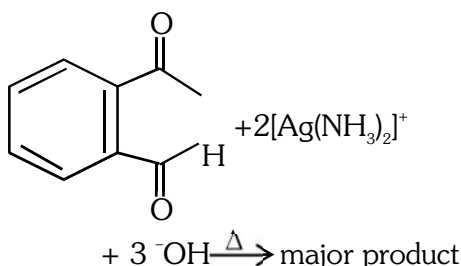
C → Pyrosulphuric acid  $\text{H}_2\text{S}_2\text{O}_7$



D → Sulphurous acid  $\text{H}_2\text{SO}_3$



**92.** Identify the major product obtained in the following reaction:



**Ans. (2)**



**93.** Pumice stone is an example of -

- (1) gel (2) solid sol  
(3) foam (4) sol

**Ans. (2)**

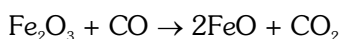
**Sol.** Pumice stone is an example of solid state

**94.** The reaction that does **NOT** take place in blast furnace between 900 K to 1500 K temperature range during extraction of iron is :

- (1)  $\text{FeO} + \text{CO} \rightarrow \text{Fe} + \text{CO}_2$   
(2)  $\text{C} + \text{CO}_2 \rightarrow 2\text{CO}$   
(3)  $\text{CaO} + \text{SiO}_2 \rightarrow \text{CaSiO}_3$   
(4)  $\text{Fe}_2\text{O}_3 + \text{CO} \rightarrow 2\text{FeO} + \text{CO}_2$

**Ans. (4)**

**Sol.** Reaction



This reaction takes place at temperature (500 K – 800 K)  
not at (900 K to 1500 K)

**95.** Which of the following statements are **INCORRECT** ?

- A. All the transition metals except scandium form MO oxides which are ionic.  
B. The highest oxidation number corresponding to the group number in transition metal oxides is attained in  $\text{Sc}_2\text{O}_3$  to  $\text{Mn}_2\text{O}_7$ .  
C. Basic character increases from  $\text{V}_2\text{O}_3$  to  $\text{V}_2\text{O}_4$  to  $\text{V}_2\text{O}_5$ .  
D.  $\text{V}_2\text{O}_4$  dissolves in acids to give  $\text{VO}_4^{3-}$  salts.  
E.  $\text{CrO}$  is basic but  $\text{Cr}_2\text{O}_3$  is amphoteric.

Choose the **correct** answer from the options given below:

- (1) B and D only                      (2) C and D only  
(3) B and C only                      (4) A and E only

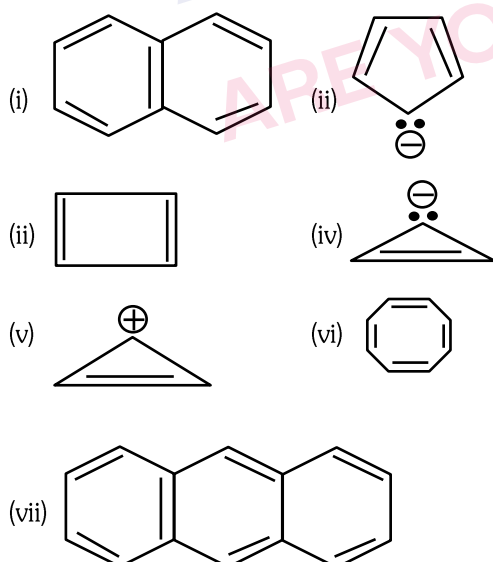
**Ans. (2)**

**Sol.**  $\text{C} \rightarrow \text{V}_2\text{O}_3 \xrightarrow{+3} \text{V}_2\text{O}_4 \xrightarrow{+4} \text{V}_2\text{O}_5 \xrightarrow{+5}$   
Acidic Nature  $\uparrow$

$\text{D} \rightarrow \text{V}_2\text{O}_5$  dissolve in acid to give  $\text{VO}_4^{3-}$  salts

This doesn't shown by  $\text{V}_2\text{O}_4$

**96.** Consider the following compounds/species:



The number of compounds/species which obey Huckel's rule is \_\_\_\_.

- (1) 6                      (2) 2                      (3) 5                      (4) 4

**Ans. (4)**

**Sol.** Huckle's rule =  $(4n + 2)\pi$  electrons

Comp (i), (ii), (v), (vii) obey Huckle's rule

**97.** What fraction of one edge centred octahedral void lies in one unit cell of fcc?

- (1)  $\frac{1}{3}$                       (2)  $\frac{1}{4}$                       (3)  $\frac{1}{12}$                       (4)  $\frac{1}{2}$

**Ans. (2)**

**Sol.**  $\rightarrow$  Edge centered octahedral void is shared between four unit cells

$\rightarrow$  Per unit cell contribution is  $\frac{1}{4}$

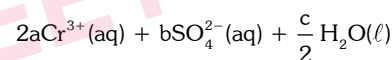
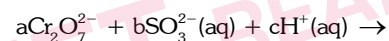
**98.** Which complex compound is most stable?

- (1)  $[\text{Co}(\text{NH}_3)_3(\text{NO}_3)_3]$   
(2)  $[\text{CoCl}_2(\text{en})_2]\text{NO}_3$   
(3)  $[\text{Co}(\text{NH}_3)_6]_2(\text{SO}_4)_3$   
(4)  $[\text{Co}(\text{NH}_3)_4(\text{H}_2\text{O})\text{Br}](\text{NO}_3)_2$

**Ans. (2)**

**Sol.** due to Chelation effect of (en).

**99.** On balancing the given redox reaction,



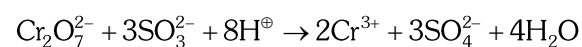
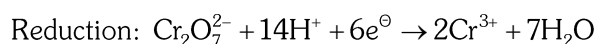
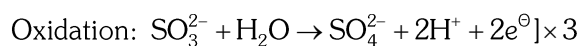
the coefficients a, b and c are found to be, respectively -

- (1) 3, 8, 1                      (2) 1, 8, 3  
(3) 8, 1, 3                      (4) 1, 3, 8

**Ans. (4)**

**Sol.** Reaction has to be balanced in acidic medium

'O' atoms are balanced by adding  $\text{H}_2\text{O}$  and then H-atom is balanced by adding  $\text{H}^+$  ions and charge is balanced by  $e^\ominus$ .

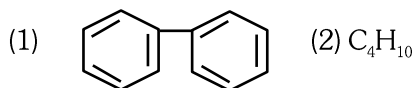
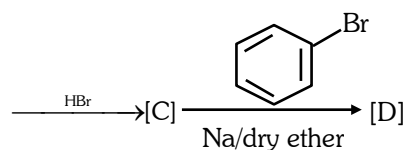
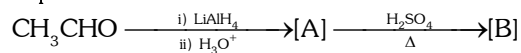


$$a = 1$$

$$b = 3$$

$$c = 8$$

**100.** Identify the final product [D] obtained in the following sequence of reactions.



**Ans. (4)**

